

maximum shifted from 400 to 435 nm, depending on diamine concentration and pH, the longer wavelength implying higher molecular weight. This condensation reaction probably occurs by hydrolysis of the starting Schiff base and subsequent formation of the more stable aromatic Schiff base.

Two main points derive from this work. First, alkyl chains were used as disposable aids to form reactive surface layers. The general criterion for spreading is that the attraction for the subphase surface be strong enough to overcome bulk cohesion.<sup>18</sup> Thus, in dialkyl terephthalaldehydes we have spreadable monomers. No polymer was obtained when terephthalaldehyde solution was used since the terephthalaldehyde quickly crystallized. After formation of the surface layer and reaction with subphase diamine, the alkylamine, being no longer needed and water-soluble, is easily removed. Second, although aliphatic Schiff bases are too unstable

to be formed in significant quantities at the air-water interface,<sup>19</sup> aromatic Schiff bases are stable. By combining these points, we produced built-up films, not obtainable by conventional LB techniques, of a polymer whose known heat resistance<sup>20</sup> is far higher than that of Langmuir-Blodgett materials used to date.

**Acknowledgment.** Thanks are due to H. Nakahara for instruction in film balance and Langmuir-Blodgett techniques.

**Registry No.** PPTA (SRU), 28157-08-6; dihexylterephthalaldehyde, 90468-36-3; dioctadecylterephthalaldehyde, 18312-95-3; (dioctadecylterephthalaldehyde)-(*p*-phenylenediamine) (copolymer), 99309-25-8; (dodecylterephthalaldehyde)-(*p*-phenylenediamine) (copolymer), 99309-26-9; dioctadecylisophthalaldehyde, 99309-24-7.

(19) Valenty, S. J. *Macromolecules* **1978**, *11*, 1221-1228.

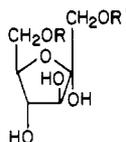
(20) 10% weight loss at 460 °C, see: Lee, H.; Stoffey, D.; Neville, K. "New Linear Polymers"; McGraw-Hill: New York, 1967; p 320.

(18) Reference 17, pp 138-139.

## Additions and Corrections

**Anomerization of Furanose Sugars and Sugar Phosphates** [*J. Am. Chem. Soc.* **1985**, *107*, 2448-2456]. JOHN PIERCE, ANTHONY S. SERIANNI, and ROBERT BARKER\*

Page 2450, Chart I: In the lower most structure (a generic *keto*-furanose), the letter R should be replaced by CH<sub>2</sub>OR. The structure would then appear as



**Hartree-Fock MO Theoretical Approach to Aromaticity. Interpretation of Hückel Resonance Energy in Terms of Kinetic Energy of  $\pi$  Electrons** [*J. Am. Chem. Soc.* **1985**, *107*, 1161-1165]. HIROSHI ICHIKAWA\* and YUKIKO EBISAWA

Page 1165: In the vertical axis of Figure 2, "KRE" should read "-KRE".

**Bimetallic Solvated Metal Atom Dispersed Catalysts. New Materials with Low-Temperature Catalytic Properties** [*J. Am. Chem. Soc.* **1984**, *106*, 2721-2722]. KENNETH J. KLABUNDE\* and YUZO IMIZU

Page 2721: Table I, 2.3% Co has an initial rate for 1-C<sub>4</sub> hydrogenation of 0.0029 (not 0.029).